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Schroeder ' s Figure
1.14. To put in some
numbers, we can use
the data from the
earlier problem. For

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water, the values given by Schroeder are (at $25\text{ C} = 298\text{ K}$):
 $= 2.57 \cdot 10^4 \text{ K}^{-1} \quad (20) \quad T$
 $= 4.52 \cdot 10^{10} \text{ Pa}^{-1} \quad (21)$

For one mole of water, the volume is $18.068 \cdot 10^{-6} \text{ m}^3$, so $C_P - C_V = (298) \cdot 18.068 \cdot 10^{-6} \cdot 2.57 \cdot 10^4 - 4.52 \cdot 10^{10} \cdot 18.068 \cdot 10^{-6} \quad (22) = 0.787 \text{ J K}^{-1} \quad (23) \quad C$

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